

## RELATIVE FORMULA MASS

You will need to refer to a periodic table to find the relative atomic masses of elements.

For each example below:

- i) state the number of atoms of each element present in the compound
- ii) calculate the relative formula mass (RFM) of the compound.

Formula of compound	Calculation of the relative formula mass
$\text{Li}_2\text{O}$	No. of atoms present: Li                      O
	RFM:
$\text{Na}_2\text{SO}_4$	No. of atoms present: Na                      S                      O
	RFM:
$\text{HNO}_3$	No. of atoms present: H                      N                      O
	RFM:
$\text{Zn}(\text{NO}_3)_2$	No. of atoms present: Zn                      N                      O
	RFM:
$\text{Mg}(\text{OH})_2$	No. of atoms present: Mg                      O                      H
	RFM:
$\text{K}_2\text{Cr}_2\text{O}_7$	No. of atoms present: K                      Cr                      O
	RFM:
$(\text{NH}_4)_2\text{SO}_4$	No. of atoms present: N                      H                      S                      O
	RFM: